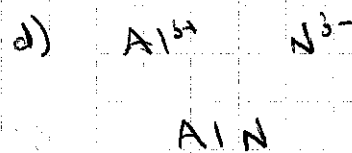
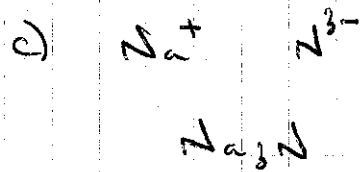
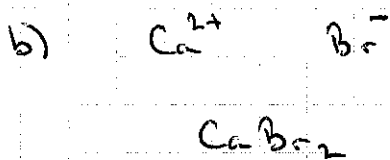
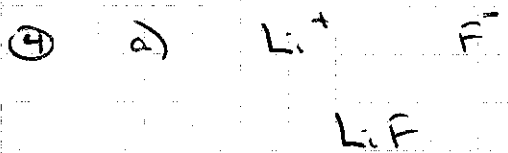
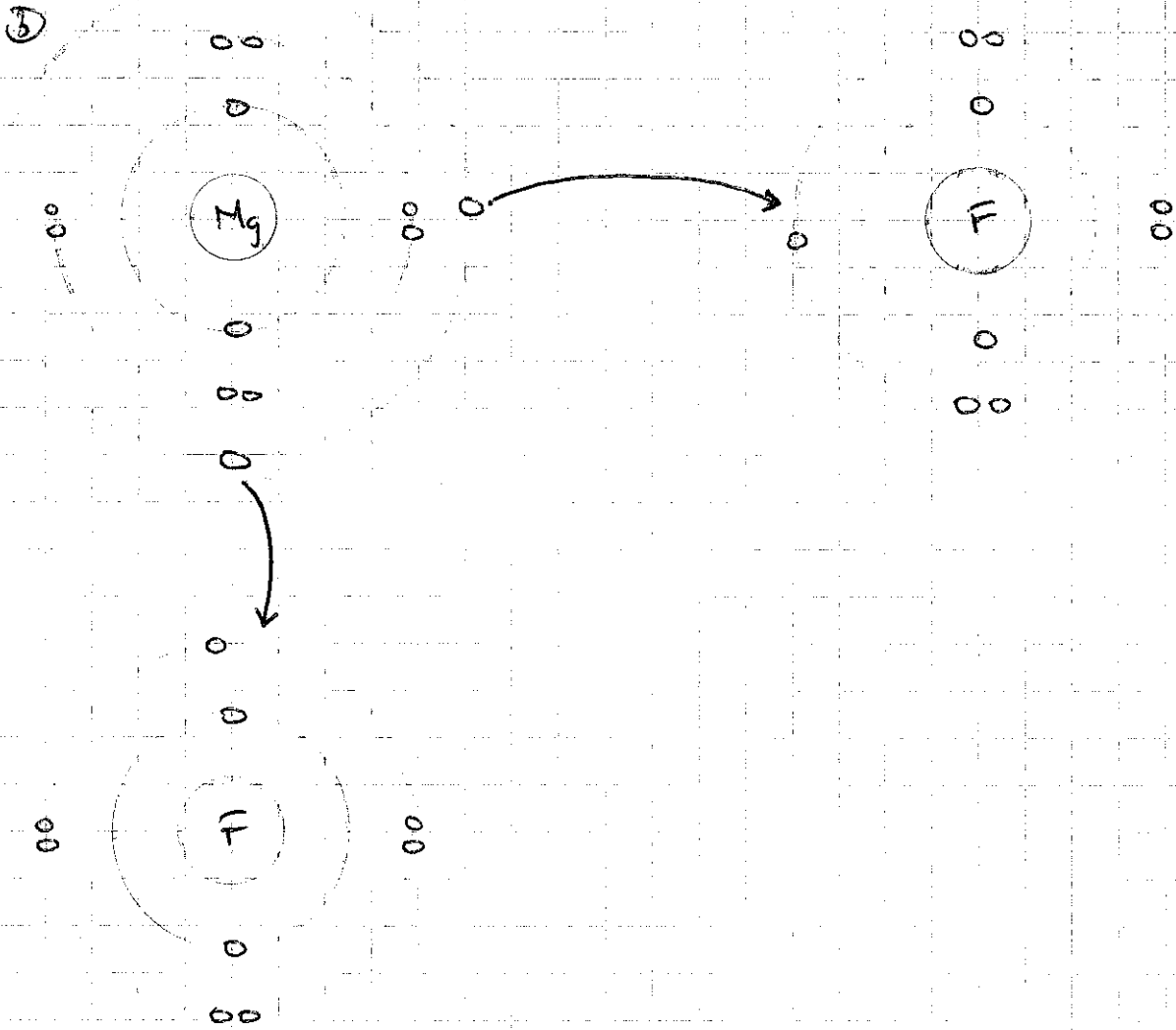


① They are the same (but opposite signs).

$$\begin{aligned} \textcircled{2} \quad \text{Al}_2\text{O}_3 &= 2\text{Al}^{3+} + 3\text{O}^{2-} \\ &= 2(+3) + 3(-2) \\ &= 0 \end{aligned}$$



- 6) a) lithium fluoride  
b) calcium bromide  
c) sodium nitride  
d) aluminum nitride

- 7) a) NaI  
b)  $\text{BeF}_2$   
c)  $\text{MgO}$   
d)  $\text{Al}_2\text{S}_3$

- 8) a) potassium chloride  
b) sodium phosphide  
c) calcium fluoride

- 9) a)  $\text{CuBr}$   
b)  $\text{CuBr}_2$   
c)  $\text{FeS}$

- 10) a) tin(II) chloride  
b) tin(IV) chloride  
c) lead(II) bromide

- 11) a)  $\text{Fe}_2\text{O}_3$   
b)  $\text{CaF}_2$   
c)  $\text{Cu}_2\text{S}$

12) iron

13) an ion that contains more than one type of element

- 14) a)  $\text{Na}_3\text{PO}_4$   
b)  $\text{CaSO}_4$   
c)  $\text{KClO}_3$   
d)  $\text{Al(OH)}_3$   
e)  $\text{Be(NO}_3)_2$   
f)  $\text{Mg(HCO}_3)_2$   
g)  $\text{NiCO}_3$

- 15) a) potassium carbonate  
b) sodium sulfate  
c) aluminum bicarbonate  
d) silver nitrate

15) b/c  $\text{NH}_4$  and  $\text{NO}_3$  are polyatomic ions and can't be separated.

16) a)  $\text{NH}_4\text{Cl}$

b)  $(\text{NH}_4)_2\text{SO}_4$